

Name: CALVIN

Date:

Hour:

Chemistry Ch.2 Review

reactant

mixture

phase

heterogeneous

product

1. starting substance in a chemical reaction
2. a physical blend of two or more components
3. part of a sample having uniform composition and properties
4. not uniform in composition
5. a substance formed in a chemical reaction

mass

homogeneous

reaction

element

compound

6. amount of matter an object contains
7. describes mixture with a uniform composition
8. process in which substances are changed into different substances
9. substance that cannot be changed into simpler substances by chemical means
10. composed of two or more substances chemically combined in a fixed proportion

- ___ 11. What is matter? *anything that has mass + takes up space*
- ___ 12. A golf ball has more mass than a tennis ball because it ____. *has more matter*
- ___ 13. List TWO physical properties and ONE chemical property:

Reasonable...

- ___ A 14. A vapor is which state of matter? *gas*
- ___ 15. All of the following are physical properties of a substance in the liquid state EXCEPT ____.
- a. indefinite volume c. not easily compressed
- b. definite mass d. indefinite shape

___ 16. Give ONE physical change YOU used yesterday: *Reasonable*

___ 17. Give ONE chemical change YOU used yesterday: *reasonable*

___ 18. BRIEFLY explain the difference between homogeneous and heterogenous mixtures. Give ONE example of each:

1 phase ←

> 1 phase

reasonable

- ___ 19. Any change of state (melting, boiling, freezing, condensation, etc...) is a physical change.
- ___ 20. Give the NAME and/or formula for TWO compounds you have at your house.

Reasonable...

___ 21. BRIEFLY explain the difference between an element and compound: *2 or more types of atoms*

___ 22. What is an INTENSIVE property? Give ONE example: *1 type of atom*

___ 23. Label the reactants and products for this reaction: *↳ depends on type of matter*

Hydrogen gas + Oxygen gas → Water

reactants

product

*ex // density
color
shape
melting point
etc...*

- A 24. What must occur for a change to be a chemical reaction?
- There must be a change in chemical properties.
 - There must be a change in physical properties.
 - The change must involve a change in mass.
 - The change must involve a change in volume.
- _____ 25. Suppose you had a block of silver. List THREE physical changes you could implement:
- A 26. Which of the following indicates that a chemical change has happened during cooking? *↳ melt, polish, bend, etc ...*
- The food darkens.
 - Bubbles form in boiling water.
 - Butter melts.
 - Energy is transferred from the stove to a pan.
- D 27. Which of the following is true for all chemical reactions?
- The total mass of the reactants increases.
 - The total mass of the products is greater than the total mass of the reactants.
 - The total mass of the products is less than the total mass of the reactants.
 - The total mass of the reactants equals the total mass of the products.
- _____ 28. Compare the mass of a tree branch before and after mulching: same

Mark the state T or F below. If F, CHANGE it to make it true:

- F 29. Boiling point is a ~~chemical~~ ^{physical} property.
- F 30. Adding sugar to water is a ~~chemical~~ ^{physical} change.
- F 31. Extensive properties depend on the ~~type~~ ^{amount} of matter.
- T 32. Intensive properties depend on the *type* of matter.
- T 33. One reactant in a combustion reaction is oxygen.
- T 34. Paper is a mixture with two phases (wood fibers and air).
- T 35. Flammability of gasoline is a chemical property.
- F 36. Oil is separated into components (diesel, gasoline, kerosene, etc) by ~~filtration~~ ^{distillation}.

GO VIKINGS!!!