

Ch. 8 PS#1

3) Noble gases, monatomic

4) NO N ~ 1 O ~ 1

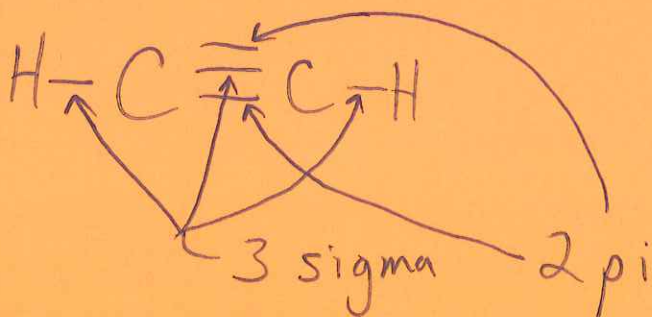
N₂O N ~ 2 O ~ 1

5) N₂ or O₂

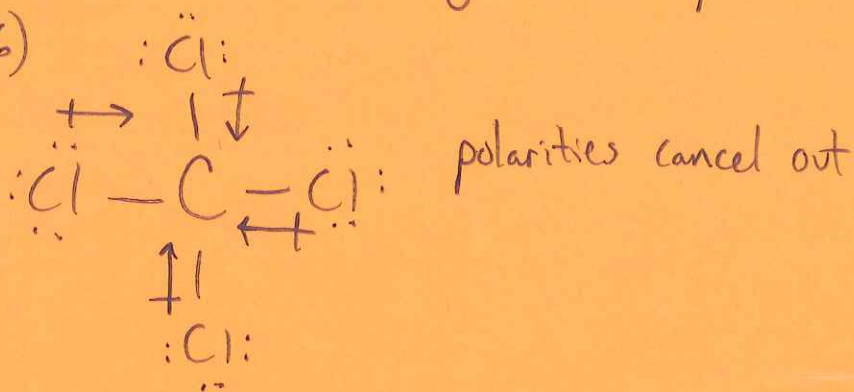
13) Noble gas configuration

24) Molecules take the shape that places valence electron pairs as far apart as possible

28)



36)

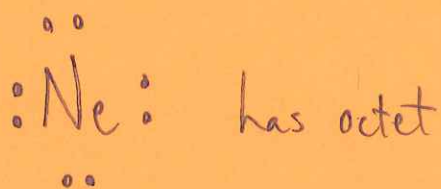


40) a) C₆H₈O₆ C ~ 6 H ~ 8 O ~ 6

b) C₁₂H₂₂O₁₁ C ~ 12 H ~ 22 O ~ 11

c) C₇H₅N₃O₆ C ~ 7 H ~ 5 N ~ 3 O ~ 6

42)



44) Ionic ~ attracted by opposite charges

Covalent ~ attracted by sharing electrons

58) Use page 177

a) H-Cl 0.9

b) H-C 0.4

c) H-F 1.9

d) H-O 1.4

e) H-H 0.0

f) S-Cl 0.5

	H	Cl	Subtract EN
←	2.1	3.0	3.0
			<u>-2.1</u>
			0.9

c, d, a, f, b, e

68) a) trigonal planar

b) trigonal pyramidal

c) linear

d) tetrahedral

GO VIKINGS! [▽]