Lab Chapter 19 directions

1. Left (RED) side of buret is for ACID.

2. Right (YELLOW) side of buret is for BASE.

3. Make sure that burets are closed. Pour to the 0.00 mL level for both burets.

4. Fill into a flask from ACID buret approximately 10 mL of ACID. Add a splash of water and 5 drops of Indicator.

5. Add BASE until the pink color appears and is maintained. Add a drop of ACID to change from pink back to colorless. Add a drop of BASE to return to pink color. This is the EQUIVALENCE point.

7. Measure and record mL's of ACID and BASE used in data table.

8. Refill burets to 0.00 mL level and rinse flask and start again for Trial 2.

9. After you have Trial 3, rinse flask, and work on calculations in class.

May 14-3:14 PM

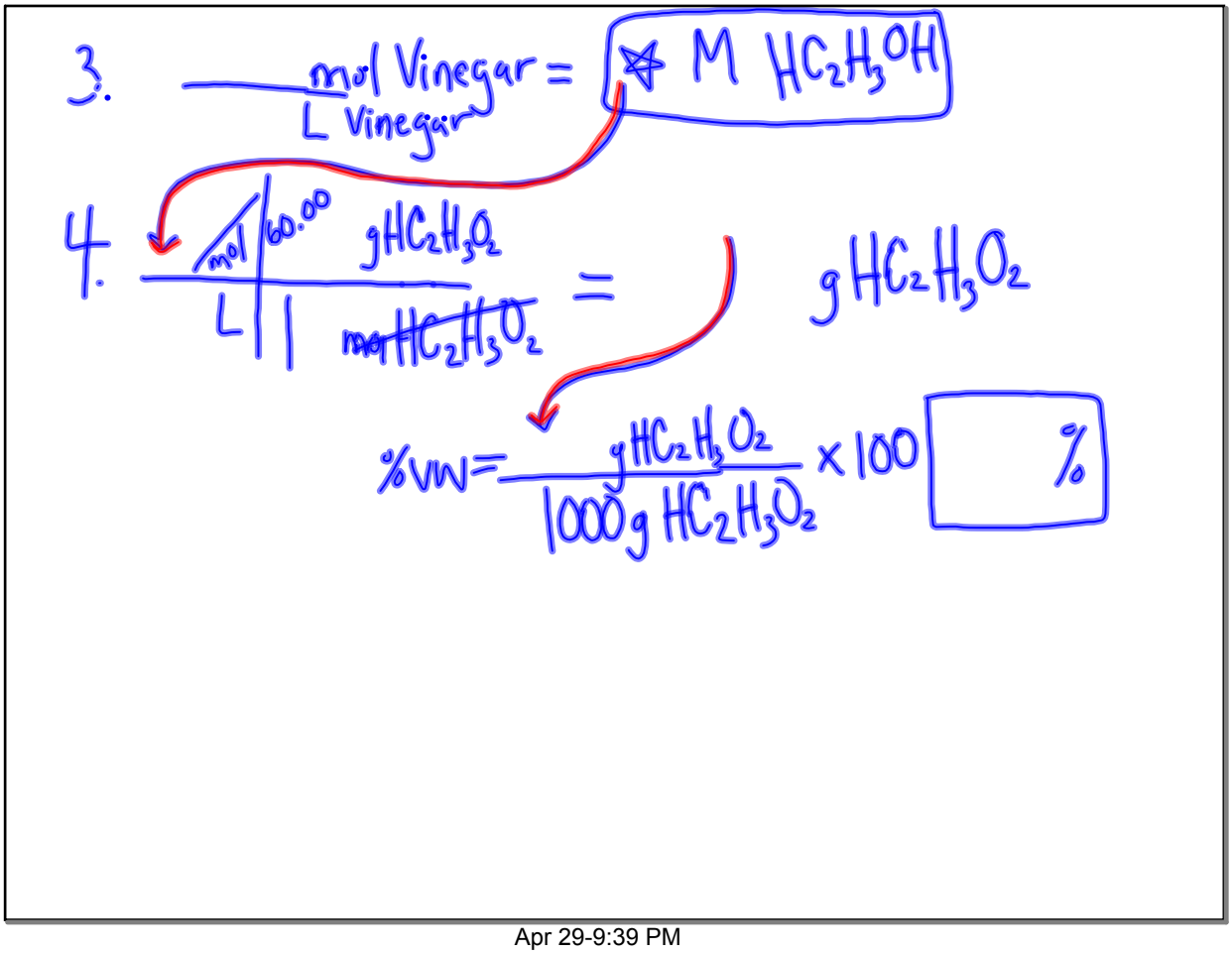
Calc

$HC_2H_3OH + NaOH \rightarrow NaC_2H_3O_2 + H_2O$

① $\frac{1\ L\ NaOH}{1\ L\ NaOH} \left| \frac{0.5\ mol\ NaOH}{} \right. = mol\ NaOH$

② $\frac{mol\ NaOH}{mol\ NaOH} \left| \frac{1\ mol\ HC_2H_3OH}{} \right. = mol\ HC_2H_3OH$

Apr 27-7:19 AM


3.
mol Vinegar / L Vinegar = ☆ M HC_2H_3OH
4.
mol / 60.00 $gHC_2H_3O_2$ / L | | $mol HC_2H_3O_2$ = $g HC_2H_3O_2$
%w/w = $\frac{g HC_2H_3O_2}{1000 g HC_2H_3O_2} \times 100$ %

Apr 29-9:39 PM